

Predicting Reactions

REDOX PRACTICE

- a) sodium metal is added to water.
- b) a solution of tin(II) chloride is added to a solution of iron(III) sulfate.
- c) an acidified solution of potassium permanganate is added to a solution of sodium sulfite.
- d) a piece of lithium metal is dropped into a container of nitrogen gas.
- e) a piece of iron is added to a solution of iron(III) sulfate.

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Practice Group 2

- a) a solution of potassium iodide is added to an acidified solution of potassium dichromate.
- b) a strip of magnesium is added to a solution of silver nitrate.
- c) solid potassium chlorate is heated in the presence of manganese dioxide as a catalyst.
- d) a stream of chlorine gas is passed through a solution of cold, dilute sodium hydroxide.
- e) solid lithium hydride is added to water.

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Practice Group 3

- a) chlorine gas is bubbled into a solution of potassium iodide.
- b) solid calcium is added to warm water.
- c) solid silver is added to a dilute nitric acid (6M) solution.
- d) hydrogen peroxide solution is added to a solution of iron(II) sulfate.
- e) propanol is burned completely in air.
- f) a solution of potassium iodide is electrolyzed.

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Practice Group 4

- a) a solution of tin(II) chloride is added to an acidified solution of potassium permanganate.
- b) carbon disulfide vapor is burned in excess oxygen.
- c) a solution of hydrogen peroxide is heated.
- d) a piece of solid bismuth is heated strongly in oxygen.
- e) a strip of copper metal is added to a concentrated solution of sulfuric acid.
- f) solid calcium hydride is added to distilled water.

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