

## Video Worksheet

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**MOLES**

1. It was the study of \_\_\_\_\_ of reacting gases that provided the first steps toward understanding the number relationship that exists between molecules when they react.
2. Avogadro said that equal volumes of different gases contained the same number of molecules as long as they were measured at the same temperature and \_\_\_\_\_.
3. By comparing masses of two equal volumes of gases measured under the same conditions, we have the ratio of the \_\_\_\_\_ of individual molecules.
4. A mole of carbon has a mass of 12 \_\_\_\_\_.
5. Moles of different substances all have different masses from one another but they all contain the \_\_\_\_\_ number of particles.
6. Atomic masses in the periodic table are all \_\_\_\_\_ masses based on assigning carbon a mass of 12.0000 units.
7. Using relative atomic masses, the equation will also tell us the \_\_\_\_\_ of reactants needed and products formed.
8. At the start of the experiment each flask contained \_\_\_\_\_ moles of hydrogen chloride.
9. At the start of the experiment each balloon contained different amounts of \_\_\_\_\_.
10. Epoxies are known for their excellent strength and \_\_\_\_\_.
11. The mass of a mole of any element is its \_\_\_\_\_ mass in grams.

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